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~~Stoichiometry - The Method of predicting amounts of products & reactants-Mass to Mole Conversions-Multiply the grams by the g/mol ratio in the reference table-Ex: 15g of C * 1/12.011 = 1.3 grams-Molar Ratio-Find the ratio of between 2 substances in a formula-Ex: CuCl + Zn ? ZnCl + Cu-The molar mass formula between Cu and Zn id 1 to 1-GFM-The mass in grams of 1 mole of any substance's ...~~

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Stoichiometry Study Guide. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. luis_dewendt. Terms in this set (14) Avogadros Number. A constant that represents the number of constituent particles (usually atoms or molecules) per mole of a given substance (6.02×10^{23})

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Chemistry Study Guide. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. Debolivia99. ... In a typical Stoichiometry problem, the given quantity is first converted to _____. Then, the mole ratio from the balanced equation is used to calculate the moles of unwanted substance. ... Chem Quiz 12.1-12.3. 20 terms ...

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, 154.21 g/mol ($6 \times 12 + 5 \times 1 = 77$: $154/77 = 2$) $\text{C}_{12}\text{H}_{10}\text{O}_2$, 222.24 g/mol ($12 \times 3 + 6 \times 1 + 2 \times 16 = 77$: $222/74 = 3$) $\text{C}_9\text{H}_{18}\text{O}_6$ 3. Empirical Formula from Percentage Composition (don't round until

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the very end!!!!) A compound is composed of 12.590% hydrogen, 37.404% aluminum, and 50.006% carbon. Please find the empirical formula:

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